

# ALKALINITY P&T KIT

Code 4533-DR-01 | Direct Reading Titrator, 0-200 ppm

 LaMotte

QUANTITY	CONTENTS	CODE
100	*Phenolphthalein Tablets	*T-2246-J
100	BCG-MR Indicator Tablets	2311A-J
60 mL	*Alkalinity Titration Reagent B	*4493DR-H
1	Test Tube, 5-10-12.9-15-20-25 mL, glass, w/cap	0608
1	Direct Reading Titrator, 0-200 Range	0382



\*Reagent is a potential health hazard. **READ SDS:** [lamotte.com](http://lamotte.com)  
**Emergency information:**  
Chem-Tel USA 1-800-255-3924  
Int'l, call collect, 813-248-0585

To order individual reagents or test kit components, use the specified code number.



Warning! This set contains chemicals that may be harmful if misused. Read cautions on individual containers carefully.  
Not to be used by children except under adult supervision.

NOTE: Read the Direct Reading Titrator Instruction Manual [1649] before proceeding. Each minor division on the Titrator scale equals 4 ppm CaCO<sub>3</sub>.

## PHENOLPHTHALEIN (P) ALKALINITY PROCEDURE

1. Fill test tube [0608] to 5 mL line with sample water.
2. Add one \*Phenolphthalein Tablet [T-2246]. Cap and mix until tablet disintegrates. If solution does not turn red, P Alkalinity is zero. If solution turns red, proceed to Step 3.
3. Fill the Direct Reading Titrator [0382] with \*Alkalinity Titration Reagent B [4493DR]. Insert Titrator into center hole of test tube cap.
4. While gently swirling tube, slowly press plunger to titrate until red color disappears. Read the test result directly from the scale where the large ring on the Titrator meets the Titrator barrel. Record as ppm P Alkalinity as CaCO<sub>3</sub>.
- EXAMPLE: Plunger tip is 3 minor divisions below line 80. The test result is 80 plus 12 (3 divisions x 4) equals 92 ppm.
5. If plunger tip reaches the bottom line on the Titrator scale [200 ppm] before the color change occurs, refill the Titrator and continue the titration. When recording test result, be sure to include the value of original amount of reagent dispensed [200 ppm].

NOTE: Do not move Titrator plunger after the P Alkalinity endpoint has been obtained. The titration is continued in the T Alkalinity procedure.

## TOTAL (T) ALKALINITY PROCEDURE

6. Being careful not to move plunger, remove Titrator and cap from test tube containing sample that was titrated in Step 4.
7. Add one BCG-MR Tablet [2311A]. Cap and shake until tablet disintegrates. Solution will turn green-blue.
8. Reinsert Titrator in cap and continue titration until color changes from green-blue to pink. Read the test result directly from the scale where the large ring on the Titrator meets the Titrator barrel. Record as ppm T Alkalinity as CaCO<sub>3</sub>. Be sure to include in test result value the total amount of titration reagent dispensed [i.e. see Step 5].
9. If only Total Alkalinity is to be tested, perform Steps 1,7, and 8 only, using a full Titrator in Step 8.

## CALCULATION OF ALKALINITY RELATIONSHIPS

Results obtained from Phenolphthalein and Total Alkalinity determinations offer a means for the stoichiometric classification of three principal forms of alkalinity present in many water supplies. Classification attributes the entire alkalinity to bicarbonate, carbonate and hydroxide; and assumes absence of other weak acids of inorganic or organic composition, such as silicic, phosphoric, and boric. This classification system further assumes incompatibility of hydroxide and bicarbonate alkalinities in the same sample. Since calculations are on a stoichiometric basis, ion concentrations in the strictest sense are not represented in the results.

According to this scheme:

- A. Carbonate Alkalinity is present when Phenolphthalein Alkalinity is not zero but less than the Total Alkalinity.
- B. Hydroxide Alkalinity is present if Phenolphthalein Alkalinity is more than one-half the Total Alkalinity.
- C. Bicarbonate Alkalinity is present if Phenolphthalein Alkalinity is less than one-half the Total Alkalinity.

Mathematical conversion of the result is shown in the following table:

Relationships Between Phenolphthalein, Total, Carbonate, Bicarbonate & Hydroxide Alkalinity			
Result of Titration	Hydroxide Alkalinity as $\text{CaCO}_3$	Carbonate Alkalinity as $\text{CaCO}_3$	Bicarbonate Alkalinity as $\text{CaCO}_3$
$P=0$	0	0	T
$P < 1/2T$	0	$2P$	$T-2P$
$P = 1/2T$	0	$2P$	0
$P > 1/2T$	$2P-T$	$2[T-P]$	0
$P=T$	T	0	0